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(21) International Application Number: PCT/FI91/00397 (22) International Filing Date: 18 December 1991 (18.12.91) (30) Priority data: 906427 28 December 1990 (28.12.90) FI (71) Applicant (for all designated States except US): NESTE OY [FI/FI]; Keilaniemi, SF-02150 Espoo (FI). (72) Inventors; and (75) Inventors/Applicants (for US only) : AHVENAINEN, Antero [FI/FI]; Jousitie 2B A2, SF-06150 Porvoo (FI). SARANTILA, Kari [FI/FI]; Paimenpojantie 1, SF-06400 Porvoo (FI). ANDTSJÖ, Henrik [FI/FI]; Jousitie 29, SF-06150 Porvoo (FI).		(74) Agent: FORSSÉN & SALOMAA OY; Georgsgatan 30, SF-00100 Helsingfors (FI). (81) Designated States: AT (European patent), BE (European patent), CH (European patent), DE (European patent), DK (European patent), ES (European patent), FR (European patent), GB (European patent), GR (European patent), IT (European patent), JP, LU (European patent), MC (European patent), NL (European patent), NO, SE (European patent), US. Published <i>With international search report.</i> <i>In English translation (filed in Finnish).</i>
(54) Title: A METHOD FOR HOMO- OR COPOLYMERIZING ETHENE (57) Abstract <p>The invention relates to a method for homo- or copolymerizing ethene in the presence of a Ziegler-Natta catalyst and a possible comonomer and hydrogen for preparing a homo- or copolymer of ethene present in a particle form in propane. In accordance with the invention, the polymerization is performed in a loop reactor in such conditions, where the temperature is higher than the critical temperature of a mixture formed by ethene, propane and a possible hydrogen and a comonomer, but lower than the melting temperature of the ethene polymer forming in the polymerization, and the pressure is higher than the critical pressure of said mixture.</p>		

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A method for homo- or copolymerizing ethene

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The invention relates to a method for homo- or copolymerizing ethene in the form of solid polymer particles.

- 10 The polymerization of ethene in the form of solid particles in a liquid phase is known. In such processes, the polymerization is typically performed in a reaction medium or a diluent, which is formed by isobutane, pentane, hexane or some other saturated aliphatic hydrocarbon. The polymerization is often performed in the presence of a Ziegler-Natta catalyst at an elevated temperature. In addition
15 to monomers, hydrogen is often used as a modifier in the polymerization, by means of which the molecular weight of the polymer to be prepared and its other properties may be affected.

- Processes are also known, in which liquid propane is used as a reaction medium.
20 For example the US patent 4 754 007 describes a process for preparing an ethene copolymer in the form of particles by using liquid propane as a liquid phase. It is postulated that certain problems related to the usage of the above-mentioned hydrocarbons may be avoided in this way. These disadvantages include for example that the usable temperature is limited, since polyethene
25 dissolves in the reaction medium at high temperatures. Therefore, the separation of comonomers and the reaction medium from each other is difficult. Furthermore, the copolymers to be prepared have a relatively wide molecular-weight distribution.

- 30 In accordance with the US patent 4 754 007, liquid propane is used as a reaction medium. In addition to the cheapness of propane, the advantages to be achieved in this way include that propane has a lower vaporization heat and it evaporates more easily than e.g. isobutane. The use of propane thus makes a more effective

separation of the reaction medium from the polymer possible. Furthermore, the use of propane helps to keep the reactor clean when polymerizing ethene in the form of particles.

- 5 The US patent especially emphasizes that the polymerization should be performed below the critical temperature of propane, which is 96.8°C. According to the patent, the process is performed in a batch reactor, in which a considerable gas volume prevails above the liquid phase. If a continuous loop reactor is preferred, the situation is completely different. The reactor has then to be
- 10 completely filled with the liquid phase, since the gas bubbles in the reaction medium would cause cavitation and thereby wearing in the circulation pumps for the reaction medium. The situation worsens, if it is desirable to prepare in the loop reactor a product with a higher melt viscosity by adding hydrogen to the reaction medium. When operating below the critical point, as is the case in the
- 15 US patent 4 754 007, separation of phases occurs as hydrogen is bubbling, and pressure shocks arise, which means that the process cannot be applied to a loop reactor.

It may be indicated that when the reaction mixture of a loop reactor contains 93

20 mol-% of propane and 7 mol-% of ethene, the critical point of the mixture is $T_c = 92.9^\circ\text{C}$, $P_c = 44.1$ bar. If the reaction medium contains 7 mol-% of ethene and 2 mol-% of hydrogen (the remainder being propane), the maximum operable temperature of the reactor is 75°C, which is far too low a temperature from the point of view of the polymerization. In the case of a mixture containing 7

25 mol-% of ethene and 2.5 mol-% of hydrogen (the remainder being propane), the maximum operable temperature is only 60°C. It is thus apparent that the process according to said US patent cannot be realized in a loop reactor.

However, it has been observed in accordance with the invention that a loop

30 reactor may be used, if the polymerization is performed above the critical point. More specifically, this relates to a critical point of a fluid mixture, which is formed by ethene acting as a monomer (+ possible comonomers), propane and

possibly hydrogen acting as a modifier. Ethene and hydrogen decrease said critical point to some extent and the critical point may be determined experimentally, when the concentrations of the components in the reaction medium are known.

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The invention thus relates to a method for homo- or copolymerizing ethene in the presence of a Ziegler-Natta catalyst and a possible comonomer and hydrogen for preparing a homo- or copolymer of ethene present in a particle form in propane. The inventive method is characterized in that the polymerization is performed in a loop reactor in such conditions, where the temperature is higher than the critical temperature of a fluid mixture formed by ethene, propane and a possible hydrogen and a comonomer, but lower than the melting temperature of the ethene polymer forming in the polymerization, and the reaction pressure is higher than the critical pressure of the mixture.

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When using a propane phase at a supercritical state, several advantages may be achieved:

1) The hydrogen content of the reactor may be adjusted within a wide range, since the operation occurs above the critical point. Then, separate liquid and gas phases can no longer be indicated, but a single phase is concerned. The separation of the phases cannot thus occur before the hydrogen concentration increases to a very high level. Because of this, more polymer types may be prepared in the reactor than when using e.g. isobutane or when using propane and operating below the critical point. Hydrogen may be added sufficiently, since the separation (bubbling) of the hydrogen or gases does not occur at the required hydrogen concentrations.

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2) Since the reactor is operated in supercritical conditions, the compressibility of the reaction fluid is good and the start-up and the operation (pressure regulation) of the reactor are facilitated, since no pressure-shock effects occur. Changes in the fluid volume always occur in the discharge of the product, which results in

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bubbling. When operating in a supercritical state, no bubbling phenomenon occurs.

3) The solubility of ethene is lower in propane than in isobutane, whereby especially polymer particles having a high melt viscosity, such as 50-500, do not dissolve in the medium and the reactor dirties less.

4) Since the boiling point of propane is low, the hydrocarbons may be separated from the polymer particles more effectively after the polymerization.

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The inventive method is adaptable for homo- or copolymerizing ethene. Suitable comonomers include e.g. propene, 1-butene, 1-pentene, 1-hexene, 4-methyl-1-pentene, 1-octene, 1-decene, 1-dodecene and 1-octadecene.

15 The usage of hydrogen as a modifier is especially well adaptable to the inventive method, when catalysts of a Ziegler-Natta type are used and when polymers with a higher melt viscosity are desired. For the above-mentioned reasons, a sufficient amount of hydrogen can always be added.

20 The polymerization is thus performed in conditions, in which the temperature is higher than the critical point of the mixture formed by propane, ethene and a possible comonomer as well as hydrogen and lower than the melting point of the ethene polymer to be produced, and the pressure is lower than the critical pressure of the reaction mixture. The polymer to be formed is then in the form of separate particles, which have an essentially uniform size and shape.

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The inventive polymerization is thus preferably performed at a temperature range 90-110°C. Since the density of the fluid in a supercritical state is strongly dependent on the pressure in the vicinity of the critical point, the operating pressure has to be selected so high that the effect of the pressure on the density is smaller. A sufficient pressure range depends on the hydrogen concentration used, but a suitable pressure in the polymerization reactor is generally 60-80 bar.

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Hydrogen may be added to the reactor up to 4 mol-% without adversely affecting the operability of the reactor. The polymerization reaction is generally performed in propane, to which ethene acting as a monomer, possible comonomers, as well as hydrogen and a catalyst are added. The polymer is removed
5 cyclically from the reactor, and the polymer is separated from propane and the monomers in a conventional manner. The polymer is preferably separated from the other substances by means of flash technique by reducing the pressure.

Any catalysts used conventionally in the polymerization of ethene may be used
10 as polymerizing catalysts. Ziegler catalysts are especially suitable to be used in the inventive method. They are catalysts, which contain transition metal components from the metals of the groups IV, V and VI of the Periodic Table of Chemical Elements. Titanium-containing catalysts are especially suitable and the catalysts used in the inventive method may also contain a support, which may be
15 e.g. silica, alumina and silica-alumina.

In addition to Ziegler catalysts, other catalyst types, such as chromium catalysts, may be preferably used in the inventive method. The selection of the catalyst essentially depends on the type of polyethene to be prepared.

20 The inventive method is not limited to only one loop reactor. It is possible and in some cases even preferable to use two loop reactors in series, whereby very high hydrogen concentrations may be used in both reactors or in either one of the two reactors and ethene polymers and/or copolymers having a wide or
25 bimodal molecular-weight distribution may thus be prepared.

For illustrating the invention, a reference is further made to the following Examples 1-3, in which the polymerization was performed in a pilot-scale loop reactor, in which the reactor diameter was 0.25 m and length 12 m. As a catalyst
30 in the Example was used a Ziegler-Natta type catalyst on a silica carrier, in which catalyst the Ti content was 4.30% by weight, the Mg content 2.80% by weight and the Al content 1.20% by weight.

The running conditions of the loop reactor and the product properties are shown in the following Table 1.

Table 1

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<u>Running conditions</u>		<u>Ex. 1</u>	<u>Ex. 2</u>	<u>Ex. 3</u>
	Temperature (°C)	98°C	98°C	98°C
10	Pressure (bar)	65 bar	65 bar	65 bar
	Catalyst feed rate (g/h)	11	11	11
	Propane feed rate (kg/h)	32	30	30
	Ethene feed rate (kg/h)	21	22	25
	Ethene concentration (mol-%)	6.5	6.4	6.4
15	Hydrogen feed rate (g/h)	43	47	57
	Hydrogen concentration (mol-%)	1.44	1.48	1.67
	Polymer production rate (kg/h)	20	20	20
<u>Product properties</u>				
	Density (kg/m ³)	980.6	982.0	983.0
20	Melt viscosity M12 (g/10 min)	67.0	86.0	104.0
	Average particle size (mm)	0.27	0.27	0.27
	Bulk density (kg/m ³)	425	422	420

25 When the test runs according to the Examples were performed, it was observed that the reactor operability was excellent in spite of the fact that the hydrogen concentration was increased for raising the melt viscosity.

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Claims

1. A method for homo- or copolymerizing ethene in the presense of a Ziegler-Natta catalyst and a possible comonomer and hydrogen for preparing in a loop reactor a homo- or copolymer of ethene present in a particle form in propane, characterized in that the polymerization is performed in a loop reactor in such conditions, where the temperature is higher than the critical temperature of a mixture formed by ethene, propane and a possible hydrogen and a comonomer, but lower than the melting temperature of the ethene polymer forming in the polymerization, and the pressure is higher than the critical pressure of said mixture .
2. A method according to Claim 1, characterized in that the polymerization temperature lies in the range 95-110°C and the pressure in the range 60-80 bar.
3. A method according to Claim 1 or 2, characterized in that the hydrogen concentration in propane is 0-4 mol-%.
4. A method according to any of the preceding Claims, characterized in that the polymerization is performed in two or more series-connected loop reactors.
5. A method according to any of the preceding Claims, characterized in that polyethene has a high melt viscosity, such as 50-500.

INTERNATIONAL SEARCH REPORT

International Application No PCT/FI 91/00397

I. CLASSIFICATION OF SUBJECT MATTER (if several classification symbols apply, indicate all) ⁶		
According to International Patent Classification (IPC) or to both National Classification and IPC		
IPC5: C 08 F 2/06, 10/02		
II. FIELDS SEARCHED		
Minimum Documentation Searched ⁷		
Classification System	Classification Symbols	
IPC5	C 08 F	
Documentation Searched other than Minimum Documentation to the Extent that such Documents are Included in Fields Searched ⁸		
SE,DK,FI,NO classes as above		
III. DOCUMENTS CONSIDERED TO BE RELEVANT⁹		
Category ¹⁰	Citation of Document, ¹¹ with indication, where appropriate, of the relevant passages ¹²	Relevant to Claim No. ¹³
Y	US, A, 4582816 (MIRO) 15 April 1986, see column 3, line 49 - column 4, line 1, example III --	1-2,4
Y	US, A, 4754007 (PULLUKAT ET AL) 28 June 1988, see column 3, line 34 - line 42 --	1-2
Y	EP, A2, 0057420 (SUMITOMO CHEMICAL INDUSTRIES LTD.) 11 August 1982, see page 14, line 13 - line 16; page 14, line 22 - line 27; page 16, line 11 - line 17 --	1,4
A	US, A, 3242150 (J.S. SCOGGIN) 22 March 1966, see column 2, line 35 - line 60 --	1-5
<p>¹⁰ Special categories of cited documents:</p> <p>"A" document defining the general state of the art which is not considered to be of particular relevance</p> <p>"E" earlier document but published on or after the international filing date</p> <p>"L" document which may throw doubts on priority claim(s) or which is cited to establish the publication date of another citation or other special reason (as specified)</p> <p>"O" document referring to an oral disclosure, use, exhibition or other means</p> <p>"P" document published prior to the international filing date but later than the priority date claimed</p> <p>"T" later document published after the international filing date or priority date and not in conflict with the application but cited to understand the principle or theory underlying the invention</p> <p>"X" document of particular relevance, the claimed invention cannot be considered novel or cannot be considered to involve an inventive step</p> <p>"Y" document of particular relevance, the claimed invention cannot be considered to involve an inventive step when the document is combined with one or more other such documents, such combination being obvious to a person skilled in the art.</p> <p>"&" document member of the same patent family</p>		
IV. CERTIFICATION		
Date of the Actual Completion of the International Search	Date of Mailing of this International Search Report	
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III. DOCUMENTS CONSIDERED TO BE RELEVANT (CONTINUED FROM THE SECOND SHEET)		
Category *	Citation of Document, with indication, where appropriate, of the relevant passages	Relevant to Claim No
A	US, A, 3324093 (C.E.ALLEMAN) 6 June 1967, see the whole document --	1-5
A	US, A, 4007321 (SCHOLZ ET AL) 8 February 1977, see column 1, line 47 - line 56 -- -----	1-5

**ANNEX TO THE INTERNATIONAL SEARCH REPORT
ON INTERNATIONAL PATENT APPLICATION NO. PCT/FI 91/00397**

This annex lists the patent family members relating to the patent documents cited in the above-mentioned international search report. The members are as contained in the Swedish Patent Office EDP file on 28/02/92. The Swedish Patent Office is in no way liable for these particulars which are merely given for the purpose of information.

Patent document cited in search report	Publication date	Patent family member(s)	Publication date
US-A- 4582816	86-04-15	NONE	
US-A- 4754007	88-06-28	CA-A- 1256246 EP-A- 0161060	89-06-20 85-11-13
EP-A2- 0057420	82-08-11	CA-A- 1162700 GB-A-B- 2094810 JP-A- 57126805 JP-A- 57126808	84-02-21 82-09-22 82-08-06 82-08-06
US-A- 3242150	66-03-22	DE-A- 1520461 GB-A- 899156	71-09-09 00-00-00
US-A- 3324093	67-06-06	NONE	
US-A- 4007321	77-02-08	AT-B- 332118 BE-A- 826072 DE-A-C- 2409839 FR-A-B- 2262672 GB-A- 1488441 JP-A- 50142688	76-09-10 75-08-27 75-09-11 75-09-26 77-10-12 75-11-17